



Low Temperature Epitaxial Growth of High Permittivity Rutile TiO₂ on SnO₂

Hongtao Wang,^{a,z} Sheng Xu,^b and Roy G. Gordon^b

^aInstitute of Applied Mechanics, Zhejiang University, Hangzhou 310027, China

^bDepartment of Chemistry and Chemical Biology, Harvard University, Cambridge, Massachusetts 02138, USA

Thin films of high dielectric constant ($\kappa \sim 68$) rutile phase titanium dioxide (TiO₂) were grown epitaxially on tin dioxide (SnO₂) substrates, which are a low cost, more abundant alternative to ruthenium electrodes used previously. Atomic layer deposition at low temperature (250°C) was used with titanium(IV) tetrakis(isopropoxide) and hydrogen peroxide (H₂O₂) as precursors. The rutile TiO₂ thin films have crystalline grains that match the structure and orientation of the grains in the polycrystalline rutile phase SnO₂ substrates. The epitaxial relations can be clearly identified from the continuous lattice fringes across the interfaces.
© 2010 The Electrochemical Society. [DOI: 10.1149/1.3457485] All rights reserved.

Manuscript submitted May 5, 2010; revised manuscript received June 7, 2010. Published June 29, 2010.

Metal-insulator-metal (MIM) structures with high capacitance density are increasingly needed for various applications in microelectronics,¹ capacitive energy storage,² and wireless communication with integrated radio-frequency circuits.³ Great effort has been put into developing high dielectric constant (κ) materials as insulators, such as HfO₂ ($\kappa \sim 25$), ZrO₂ ($\kappa \sim 25$ –30), Ta₂O₅ ($\kappa \sim 26$), and rare-earth oxides ($\kappa \sim 20$ –30).^{1,4,5} However, these oxides only have moderately high κ values, which cannot satisfy the demanding scaling of the effective oxide thickness (EOT) for the next-generation electronic devices.⁶ Barium–strontium titanate (Ba,Sr)TiO₃ (BST) thin films are also of great interest for their possible ultrahigh dielectric constant ($\kappa \sim 300$),^{1,7} but the BST dielectric constant has a strong thickness dependence and decreases to ~ 16 for a 20 nm thick film on a Ru bottom electrode.^{8,9} It is also challenging to grow BST films on non-noble metal substrates. The IBM research group found that BST was amorphous on TaSiN barrier layers, whereas high- κ is only obtained for well-crystallized materials.⁷ Recently, there has been extensive research on growing BST thin films on base-metal substrates to lower material cost.^{10–12} Postdeposition annealing (400–750°C) in oxidizing ambient is necessary for making BST with a high dielectric constant,¹² but interfacial layers are introduced due to easy oxidation of base-metal electrodes, finally leading to a high EOT, which makes BST/base-metal stacks ineffective as well as complex to make.

Compared to the aforementioned oxides, the simple binary oxide, rutile phase TiO₂, is more attractive for its high dielectric constant (up to $\kappa \sim 114$ for films with random texture), less thickness dependence, and low dielectric loss.^{13,14} However, TiO₂ thin films grown by both physical vapor deposition and chemical vapor deposition usually have anatase or amorphous structures, which only have a moderately high dielectric constant ($\kappa \sim 20$ –40).¹⁵ High temperature annealing (above 700°C) is necessary to transform them to the rutile phase, which requires high thermal budgets and causes structural integrity problems.¹⁶ Atomic layer deposition (ALD) with halide-based precursors can deposit rutile-structured TiO₂, but residual halide impurity causes the films to have high leakage currents.¹⁷ The as-deposited rutile TiO₂ films were realized only in a few systems with substrates such as oxidized noble metals (Ru and Ir) or RuO₂¹⁷ and insulating single-crystal oxides [sapphire(001) and MgO(001)].^{4,18} These substrates have lattice planes matching rutile TiO₂ that make nucleation of the rutile phase kinetically preferable to the anatase phase.

In this article, we report, for the first time, ALD of rutile TiO₂ epitaxially on conductive SnO₂, which has potential application to high density capacitors. SnO₂ layers are ideal as bottom electrodes for several reasons: (i) SnO₂ has the rutile structure and small lattice

mismatches to rutile TiO₂ [$(a_{\text{TiO}_2} - a_{\text{SnO}_2})/a_{\text{SnO}_2} = -3.10\%$, $(c_{\text{TiO}_2} - c_{\text{SnO}_2})/c_{\text{SnO}_2} = -7.16\%$]; (ii) SnO₂ thin films can be made highly conductive by doping with fluorine (F) or antimony (Sb) (<1 mΩ cm); (iii) SnO₂ is stable under the oxidizing annealing ambients needed for low leakage through high- κ insulators; and (iv) SnO₂ has low material cost compared to noble metal electrodes such as ruthenium or iridium.

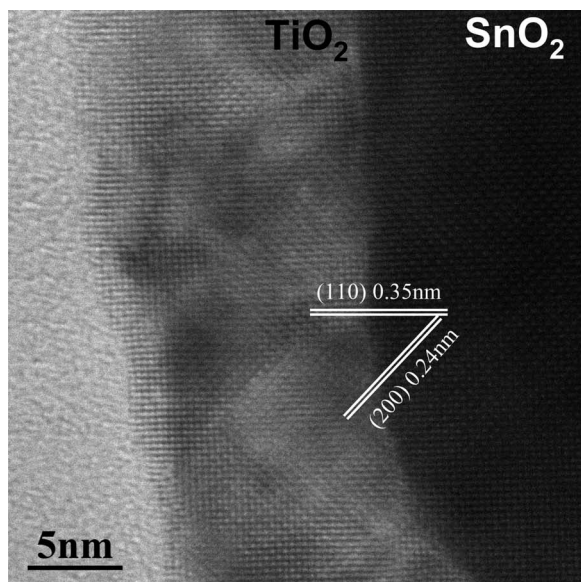
Experimental

TiO₂/SnO₂ stacks were deposited in a flow-type ALD reactor operated at a base pressure of 0.3 Torr using N₂ as purging gas. The precursors are titanium(IV) tetrakis(isopropoxide) (TTIP) and hydrogen peroxide (H₂O₂) for TiO₂, and tetramethyl tin [Sn(Me)₄] and nitrogen dioxide (NO₂) for SnO₂.¹⁹ The SnO₂ films were Sb-doped with antimony(III) ethoxide [Sb(OEt)₃].^{20,21} All precursors are commercially available (e.g., Sigma-Aldrich) and inexpensive. The bubbler temperatures were 40 and 50°C for TTIP and Sb(OEt)₃, respectively. All other precursors were kept at room temperature. TiO₂ was deposited at a substrate temperature of 250°C by alternatively dosing the vapors from TTIP and H₂O₂. A trimethylaluminum dose was inserted every 90 cycles of TiO₂ to provide aluminum doping to decrease the leakage current. The conductive SnO₂ films were deposited by repeatedly growing 20 cycles of SnO₂ followed by one dose of Sb(OEt)₃ at a substrate temperature of 400°C. The growth rates are 0.3 and 1.2 Å/cycle for TiO₂ and SnO₂, respectively, independent of the substrate. The TiO₂/SnO₂ stack structure was characterized by transmission electron microscopy (TEM) and selected area diffraction (SAD). Capacitors for the measurement of the dielectric constant were made by evaporating palladium (Pd) electrodes and patterning capacitor areas lithographically by lift-off.

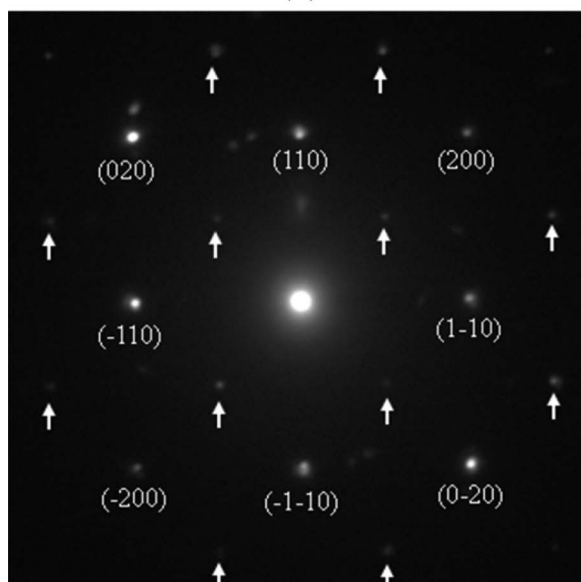
Results and Discussion

High resolution TEM was used to examine the epitaxial growth of rutile TiO₂ on an E-beam-evaporated SnO₂ substrate with grain size ~ 100 nm for the convenience of TEM imaging (Fig. 1a). The mass difference gives a clear contrast between TiO₂ (light color) and SnO₂ (dark color). The lattice fringes extend through the whole stack and are continuous across the interface. The atomic registry clearly shows that rutile TiO₂ grows epitaxially on SnO₂ at the low substrate temperature (250°C). Two lattice spacings, 0.35 and 0.24 nm, can be directly measured from the image, which are close to those of SnO₂(110) and SnO₂(200) planes. No difference in lattice spacing between TiO₂ and SnO₂ is observed in the high resolution image. The SAD pattern (Fig. 1b) has a fourfold symmetry and reveals the zone axis to be [001]. The major spots come from the SnO₂ crystal and are labeled according to the symmetry. The weak spots, indicated by arrows, distributed periodically around the major

^z E-mail: htw@zju.edu.cn



(a)



(b)

Figure 1. (a) High resolution cross-sectional TEM image of $\text{TiO}_2/\text{SnO}_2$ stack. The SnO_2 layer (200 nm) was deposited by electron-beam (E-beam) evaporation from bulk SnO_2 source followed by annealing in air for 1 h at 800°C . (b) The corresponding SAD pattern.

spots, come from the TiO_2 crystal and show the epitaxy relation to be $(110)\text{TiO}_2\parallel(110)\text{SnO}_2$; $[001]\text{TiO}_2\parallel[001]\text{SnO}_2$. Some random spots come from nearby grains due to the SAD aperture being larger than the grains and possibly also to astigmatism in the TEM.

The growth behavior of TiO_2 on conductive ALD SnO_2 is more interesting for applications. The as-deposited Sb-doped SnO_2 thin films are polycrystalline with grain size ~ 20 nm (Fig. 2). The inset SAD pattern has sharp and continuous rings, which indicate that the film has good crystallinity and random texture. The plan-view TEM image of a $\text{TiO}_2/\text{SnO}_2$ stack (Fig. 3a) shows that moiré fringes fully extend throughout each identifiable grain, which implies that both TiO_2 and SnO_2 films have comparable grain size. The corresponding SAD pattern is shown in Fig. 3b with $\text{Si}(200)$ diffraction spots intentionally recorded as an internal scale for accurately measuring the

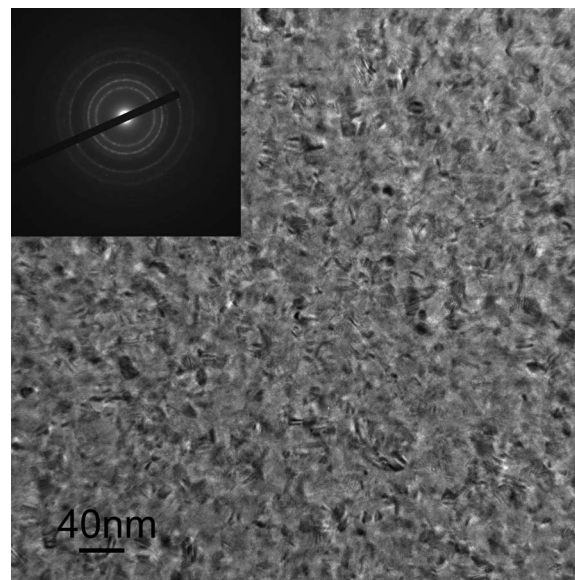


Figure 2. Plan-view TEM image of a 25 nm SnO_2 film on SiN_x membrane (50 nm) TEM grid. The inset is the SAD pattern.

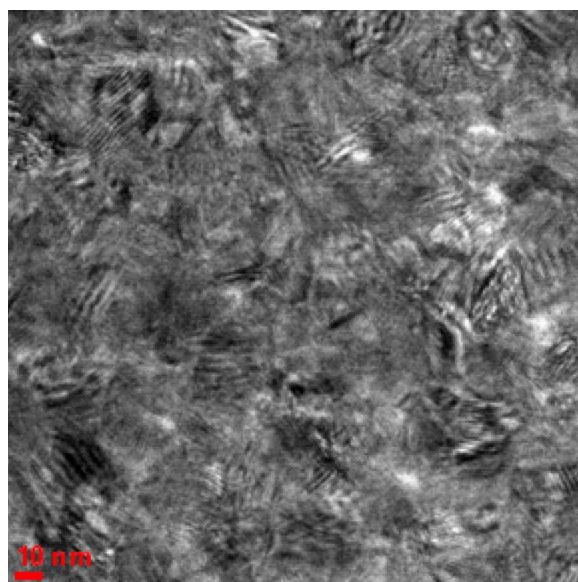
lattice spacings. All rings can be unambiguously assigned to rutile phase SnO_2 and TiO_2 by comparing the lattice spacings to the powder X-ray diffraction database (PCPDF file no. 87-0920 and no. 88-0287). The anatase phase can be excluded by the absence of a diffraction ring from its (101) plane (3.51 Å, PCPDF file no. 89-4203), which has the maximum intensity and can be clearly distinguished from rutile $\text{TiO}_2(110)$ (3.25 Å). Figure 4 confirms the local epitaxial growth of TiO_2 on the polycrystalline ALD SnO_2 layer, which also rules out the possible presence of anatase TiO_2 because of its large lattice mismatch with SnO_2 . Similar SAD measurements in many parts of several samples ruled out the presence of any significant amount of anatase phase in the films.

Capacitors with the MIM structure were fabricated to characterize the electrical properties of the ALD $\text{TiO}_2/\text{SnO}_2$ stacks. Before the top Pd electrode evaporation, the stack was annealed at 300°C for 1 h in a flow of a mixture of O_2 and O_3 (concentration 5 g/N m^3). Ozone annealing is effective in filling the oxygen vacancies and lowering the leakage current.²² Figure 5 shows the capacitance–voltage curve measured at 1 MHz for an 18 nm TiO_2 film. There is $\sim 8\%$ difference in capacitance between the applied negative and positive biases, which corresponds to the inversion and accumulation conditions of the Sb-doped SnO_2 layer, respectively. The equivalent SiO_2 thickness (EOT) is 0.98–1.07 nm and the dielectric constant is calculated to be 65–72, which is comparable to the dielectric constant of rutile TiO_2 deposited on oxidized ruthenium electrodes.²³ This high value of the dielectric constant corroborates the rutile structure of the TiO_2 films and the absence of the anatase phase, which would have a lower dielectric constant.

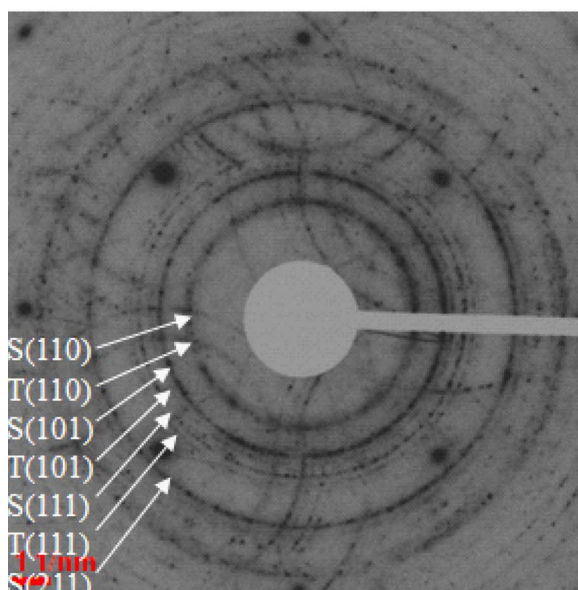
TiO_2 films grown under the same ALD conditions on other substrates, such as silicon and silicon dioxide, had the anatase crystal structure and a dielectric constant less than 40 (data not shown).

Conclusions

Rutile phase TiO_2 thin films were deposited by ALD on SnO_2 substrates at 250°C using TTIP and H_2O_2 . The TiO_2 thin films have crystallinity matching that of the SnO_2 substrates. The epitaxial relations can be clearly identified from the lattice fringes that are continuous across the interfaces. High density (3.2–3.5 $\mu\text{F}/\text{cm}^2$) MIM capacitors were fabricated on planar substrates based on ALD



(a)



(b)

Figure 3. (Color online) (a) Plan-view TEM image of an 18 nm TiO_2 /20 nm SnO_2 stack on Si substrate. (b) The SAD pattern. The discrete spots come from the electron diffraction by the Si substrate, which are doubly scattered by the polycrystalline film stack to form less intensive rings around each spot. In the label, S and T indicate rutile phase SnO_2 and TiO_2 , respectively.

rutile TiO_2 / SnO_2 stacks. The low thermal budget and the low material cost recommend ALD TiO_2 / SnO_2 stacks as potential candidates for application to high density MIM capacitors.

Acknowledgments

This work was performed in part at Harvard University's Center for Nanoscale Systems, a member of the National Nanotechnology Infrastructure Network, supported by the U.S. National Science Foundation under award no. ECS-0335765. One of the authors acknowledges the support of the Science Foundation of Chinese University (grant no. 2009QNA4034).

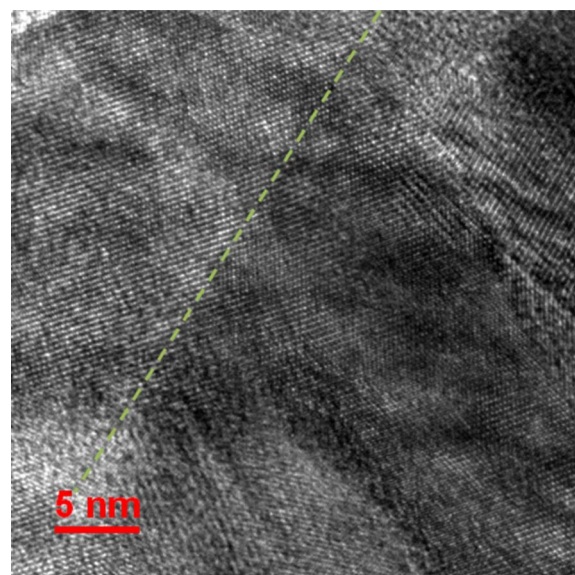


Figure 4. (Color online) Cross-sectional TEM image of a TiO_2 / SnO_2 stack. The dashed lines indicate the interface.

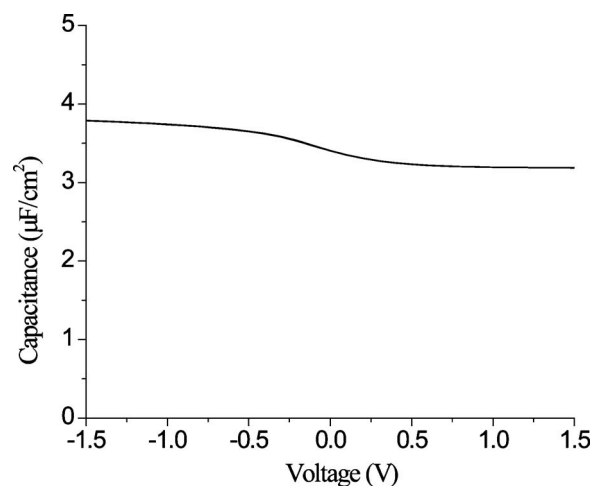


Figure 5. The capacitance density vs the applied voltage for the stack of Pd/TiO_2 (18 nm)/ SnO_2 .

References

- G. D. Wilk, R. M. Wallace, and J. M. Anthony, *J. Appl. Phys.*, **89**, 5243 (2001).
- P. Banerjee, I. Perez, L. Henn-Lecordier, S. B. Lee, and G. W. Rubloff, *Nat. Nanotechnol.*, **4**, 292 (2009).
- S.-J. Ding, J. Xu, Y. Huang, Q.-Q. Sun, D. W. Zhang, and M.-F. Li, *Appl. Phys. Lett.*, **93**, 092909 (2008).
- S. K. Kim, G. J. Choi, S. Y. Lee, M. Seo, S. W. Lee, J. H. Han, H. S. Ahn, S. Han, and C. S. Hwang, *Adv. Mater.*, **20**, 1429 (2008).
- H. T. Wang, J. J. Wang, R. Gordon, J. S. M. Lehn, H. Z. Li, D. Hong, and D. V. Shenai, *Electrochem. Solid-State Lett.*, **12**, G13 (2009).
- L. Jensen, P.-O. Astrand, and K. V. Mikkelsen, *J. Phys. Chem. B*, **108**, 8226 (2004).
- D. E. Kotecki, J. D. Baniecki, H. Shen, R. B. Laibowitz, K. L. Saenger, J. J. Lian, T. M. Shaw, S. D. Athavale, C. Cabral, P. R. Duncombe, et al., *IBM J. Res. Dev.*, **43**, 367 (1999).
- J.-H. Ahn, J.-Y. Kim, S.-W. Kang, J.-H. Kim, and J.-S. Roh, *Appl. Phys. Lett.*, **91**, 062910 (2007).
- H. Tabata and T. Kawai, *Appl. Phys. Lett.*, **70**, 321 (1997).
- P. Gonon and F. El Kamel, *J. Appl. Phys.*, **101**, 073901 (2007).
- B. Laughlin, J. Ihlefeld, and J. P. Maria, *J. Am. Ceram. Soc.*, **88**, 2652 (2005).
- A. V. Polotai, I. Fujii, D. P. Shay, G. Y. Yang, E. C. Dickey, and C. A. Randall, *J. Am. Ceram. Soc.*, **91**, 2540 (2008).
- H. P. R. Frederikse, *Permittivity (Dielectric Constant) of Inorganic Solids: CRC Handbook of Chemistry and Physics*, CRC Press, Boca Raton, FL (2008).

14. R. A. Parker and J. H. Wasilik, *Phys. Rev.*, **120**, 1631 (1960).
15. S. A. Campbell, H. S. Kim, D. C. Gilmer, B. He, T. Ma, and W. L. Gladfelter, *IBM J. Res. Dev.*, **43**, 383 (1999).
16. D. J. Won, C. H. Wang, H. K. Jang, and D. J. Choi, *Appl. Phys. A: Mater. Sci. Process.*, **73**, 595 (2001).
17. K. Fröhlich, M. Ľapajna, A. Rosová, E. Dobročka, K. Hušková, J. Aarik, and A. Aidla, *Electrochem. Solid-State Lett.*, **11**, G19 (2008).
18. M. Schuisky, A. Harsta, A. Aidla, K. Kukli, A. A. Kiisler, and J. Aarik, *J. Electrochem. Soc.*, **147**, 3319 (2000).
19. V. E. Drozd and V. B. Aleskovski, *Appl. Surf. Sci.*, **82-83**, 591 (1994).
20. A. G. Zawadzki, C. J. Giunta, and R. G. Gordon, *J. Phys. Chem.*, **96**, 5364 (1992).
21. J. Proscia and R. G. Gordon, *Thin Solid Films*, **214**, 175 (1992).
22. H. Shinriki, M. Nakata, Y. Nishioka, and K. Mukai, *IEEE Electron Device Lett.*, **10**, 514 (1989).
23. S. K. Kim, G. W. Hwang, W.-D. Kim, and C. S. Hwang, *Electrochem. Solid-State Lett.*, **9**, F5 (2006).